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Determine the molecular weight and stoichiometric mole and mass air-fuel ratios for the Oklahoma gas mole composition given in Table 3.4. Solution. Equation (3.2),  $\text{CH}_4 + 2\text{O}_2 + 2(3.76)\text{N}_2\text{CO}_2 + 2\text{H}_2\text{O} + 2(3.76)\text{N}_2$  (3.2) shows that there are  $2 + 2(3.76) = 9.52$  moles of air required for complete combustion of each mole of methane.

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This solutions manual has been prepared to accompany the 3rd edition of the author's Introduction to Internal Combustion Engines. At the end of many of the questions is a discussion, which is intended to provide useful supplementary information.

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